

Chemistry Chemistry Raymond Chang 11th Edition

Chemistry Chemistry Raymond Chang 11th Edition The Chemistry of Color A Colorful Journey Through the World of Pigments Have you ever wondered why the sky is blue why leaves are green or why a sunset paints the horizon with fiery hues The answer lies in the fascinating world of color and its intricate relationship with chemistry This article delves into the chemical makeup of pigments exploring how molecules absorb and reflect light to create the vibrant tapestry of colors we see around us

- 1 The Nature of Color Color is a result of our perception of light White light like sunlight is actually a mixture of all colors of the visible spectrum When this light interacts with matter certain wavelengths are absorbed while others are reflected Its the reflected wavelengths that reach our eyes stimulating specific photoreceptor cells in our retinas and creating the sensation of color
- 2 The Role of Pigments Pigments are substances that absorb specific wavelengths of light and reflect the rest This selective absorption is what gives them their characteristic color Understanding the chemical structure of pigments is key to understanding how they interact with light
- 3 Types of Pigments Organic Pigments These pigments are derived from natural sources like plants animals or minerals They often contain complex organic molecules with conjugated systems of alternating single and double bonds These systems allow for the absorption of specific wavelengths of light leading to vibrant colors Examples Chlorophyll green carotenoids orange and yellow anthocyanins red purple and blue Inorganic Pigments These pigments are based on metal compounds or oxides They typically have a more rigid structure and less complex bonding arrangements than organic pigments Examples Iron oxides red yellow brown cadmium sulfide yellow titanium dioxide white
- 4 The Chemistry of Pigments A Closer Look Chromophores These are specific groups of atoms within a pigment molecule responsible for absorbing light They often involve double bonds or lone pairs of electrons Examples Carbonyl groups C=O azo groups N=N quinoid rings Auxochromes These are groups of atoms that modify the absorption properties of chromophores They can shift the wavelength of maximum absorption increasing the intensity of the color Examples Amino groups NH_2 hydroxyl groups OH
- 5 Color Modification Techniques Pigment Mixing Mixing pigments allows us to create new colors When different pigments are mixed they absorb different wavelengths of light resulting in a new color that is the combination of the reflected wavelengths Dyeing Dyes are soluble pigments that

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